

SI units

length → meter m

mass → kilogram kg kilo = 10^3

time → seconds s

temperature → Kelvin K

Celsius ($^{\circ}\text{C}$) ~~$^{\circ}\text{C}$~~

Kelvin uses Celsius scale →
starts at absolute zero.

T Tera 10^{12}

G Giga 10^9

M Mega 10^6

kilo k 10^3

hecto h 10^2

deca da 10^1

base 10^0

Prefixes

deci d 10^{-1}

centi c 10^{-2}

milli m 10^{-3}

micro μ 10^{-6}

nano n 10^{-9}

128 km → cm

$$128 \frac{\text{km}}{\text{km}} * \frac{1000 \frac{\text{m}}{\text{m}}}{1 \frac{\text{km}}{\text{km}}} * \frac{100 \frac{\text{cm}}{\text{m}}}{1 \frac{\text{m}}{\text{m}}}$$

12,800,000 cm

$$128 \frac{\text{km}}{\text{km}} * \frac{10^3 \frac{\text{m}}{\text{m}}}{1 \frac{\text{km}}{\text{km}}} * \frac{10^2 \frac{\text{cm}}{\text{m}}}{1 \frac{\text{m}}{\text{m}}}$$

128 * 10^5 cm

1.28 * 10^7 cm

	T	Tera	10^{12}
3	G	Giga	10^9
3	M	Mega	10^6
3	K	kilo	10^3
2	H	hecto	10^2
1	d	deca	10^1
		base	10^0
2	d	deci	10^{-1}
1	c	centi	10^{-2}
0	m	milli	10^{-3}

128 km → cm

128 0 0 0 0 0
uuuu

128 0 0 0 0 0 cm

793 mg → — Mg

0 0 0 0 0 0 793 Mg
uuuuuuu

$$793 \text{ mg} * \frac{1 \text{ g}}{1000 \text{ mg}} * \frac{1 \text{ Mg}}{1000000 \text{ g}}$$

$$\cancel{793} \overline{111111111} \\ \cancel{100000000}$$

$$\text{deci} \rightarrow 10^{-1} d$$

$$\text{kilo} \rightarrow 10^3 k$$

$$\text{mega} \rightarrow 10^6 M$$

$$\text{centi} \rightarrow 10^{-2} c$$

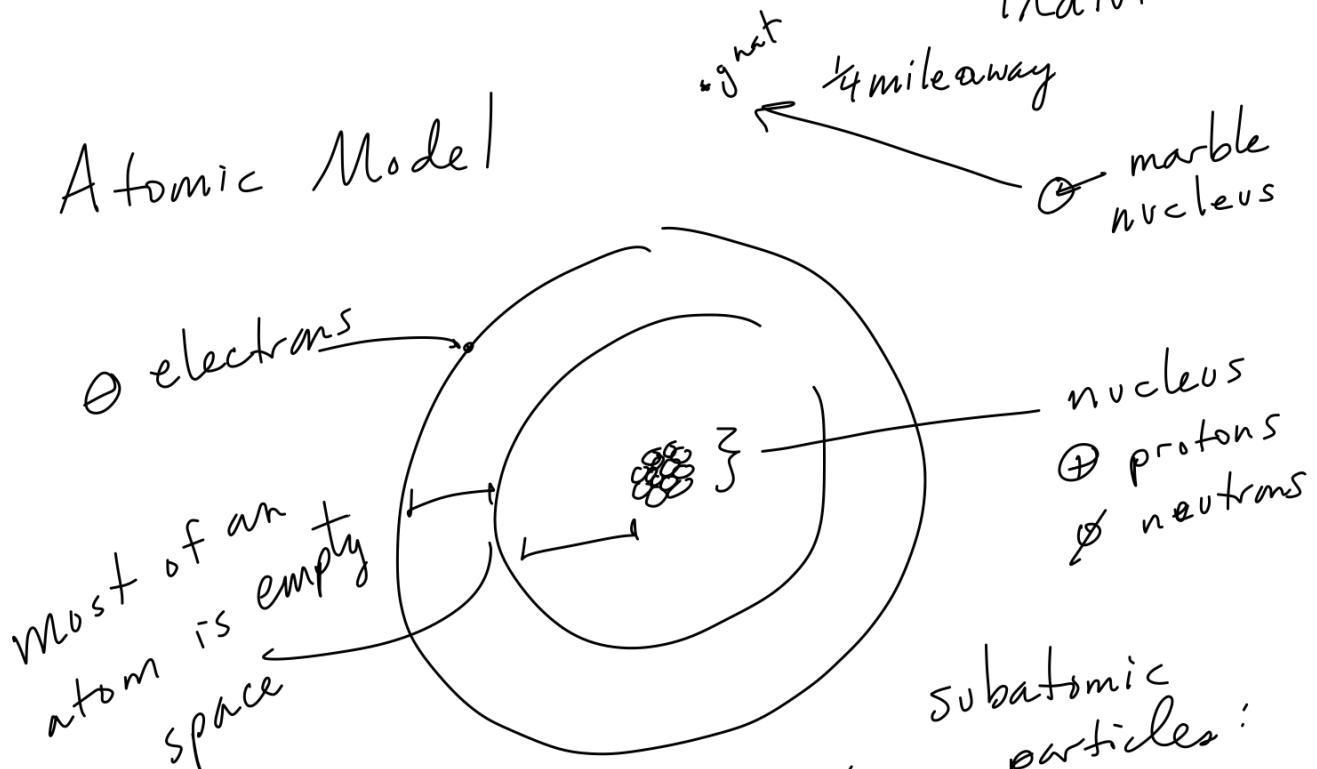
$$\text{micro} \rightarrow 10^{-6} \mu$$

$$\text{nano} \rightarrow 10^{-9} n$$

$$\text{giga} \rightarrow 10^9 G$$

Democritus → cut in half → 'atoms'
indivisible

Atomic Model



quarks protons arrangement
^{up, down,}
^{top, bottom,}
^{strange, charmed}

neutrons of 3 quarks

subatomic particles:

protons \oplus
 neutrons \ominus
 electrons \ominus

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$D = \frac{M}{V}$$

$$V(D) = \left(\frac{M}{V}\right)V$$

mass 208 g

volume 8.0 L

$$M = VD$$

$$D = \frac{208 \text{ g}}{8.0 \text{ L}} = \boxed{26 \text{ g/L}}$$

$$D = \frac{M}{V}$$

mass = ?

$$\text{volume} = 42 \text{ mL}$$

$$\text{density} = 5.5 \text{ g/mL}$$

$$V(D) = \left(\frac{M}{V}\right)V$$

$$M = VD = (42 \text{ mL})(5.5 \text{ g/mL}) \\ = \boxed{231 \text{ g}}$$

Volume = ?

$$\text{mass} = 24 \text{ g}$$

$$\text{density} = 6.0 \text{ g/mL}$$

$$D = \frac{M}{V}$$

$$V = \frac{M}{D} = \frac{24 \text{ g}}{6.0 \text{ g/mL}}$$

$$\boxed{4.0 \text{ mL}}$$

HW
make sure you
understand
outline
online HW WK 2
due Sep 30th
~~Quiz 1~~
due Sep 23rd
Quiz #1