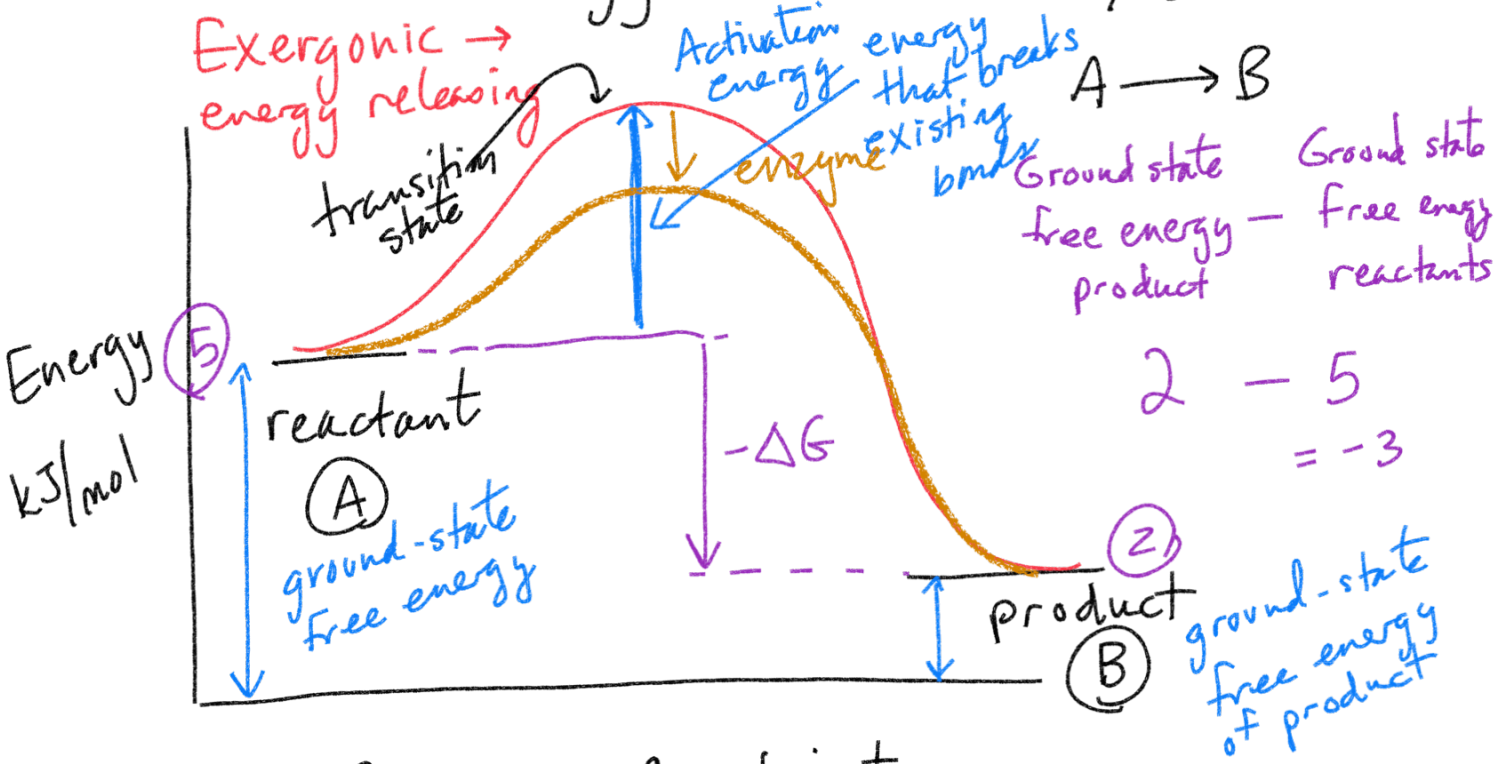
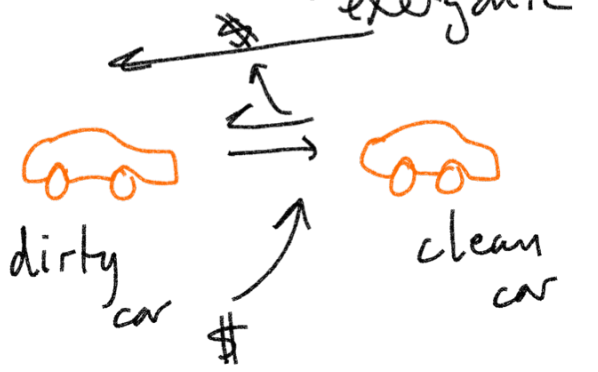
W-GB General Biology Week 21 2/22



Favorable spontaneous

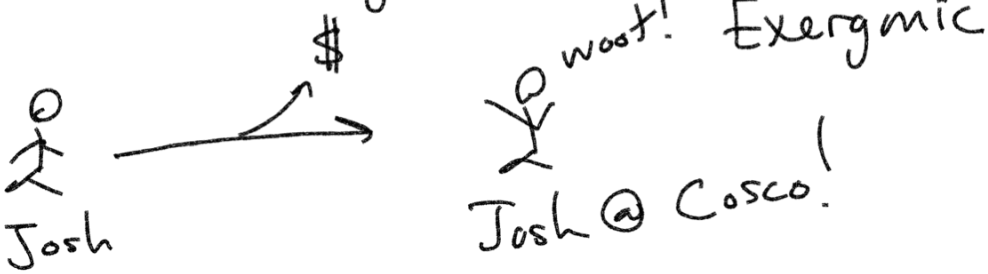
Nature favors low energy = stability

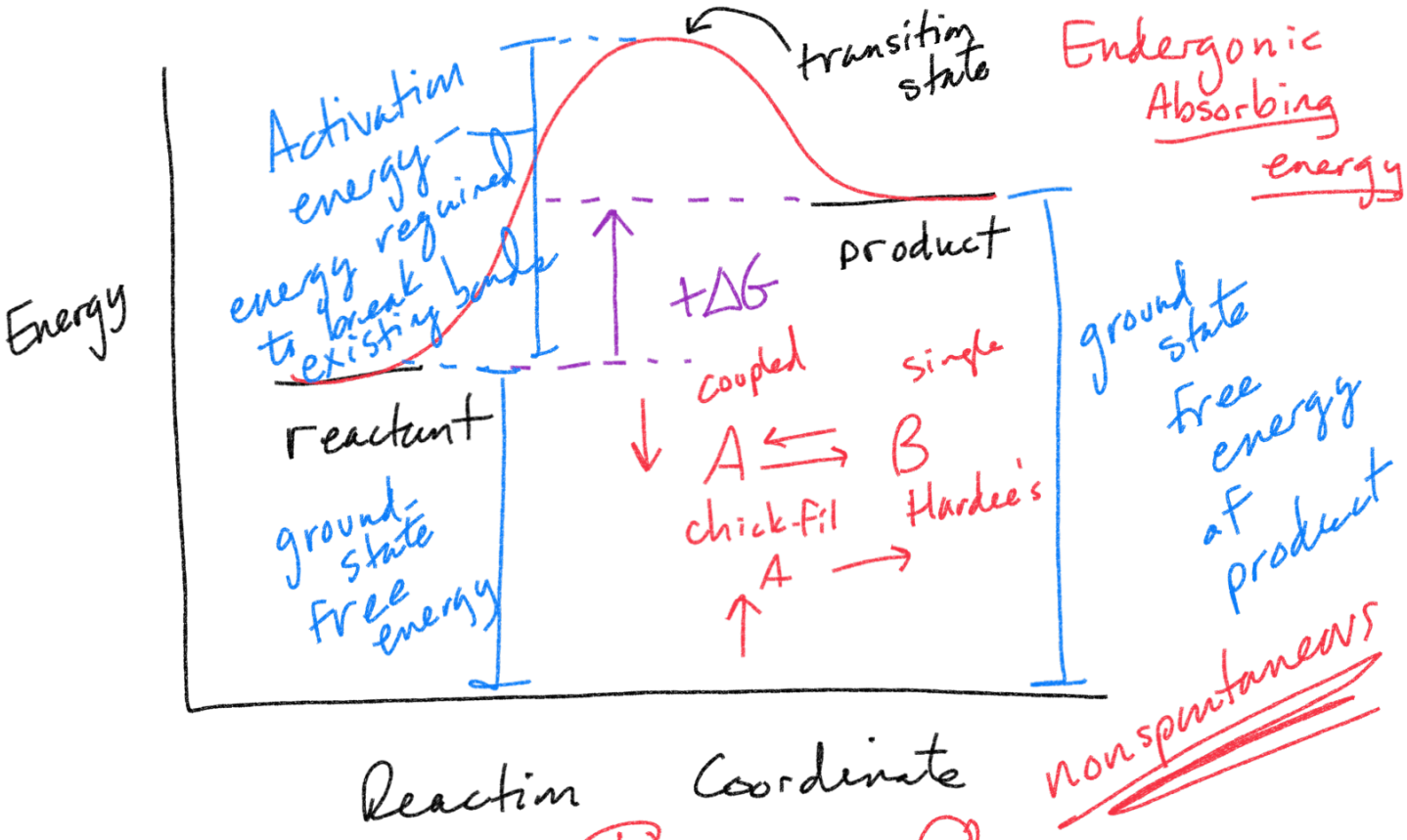
exergonic



Endergmic

endergonic



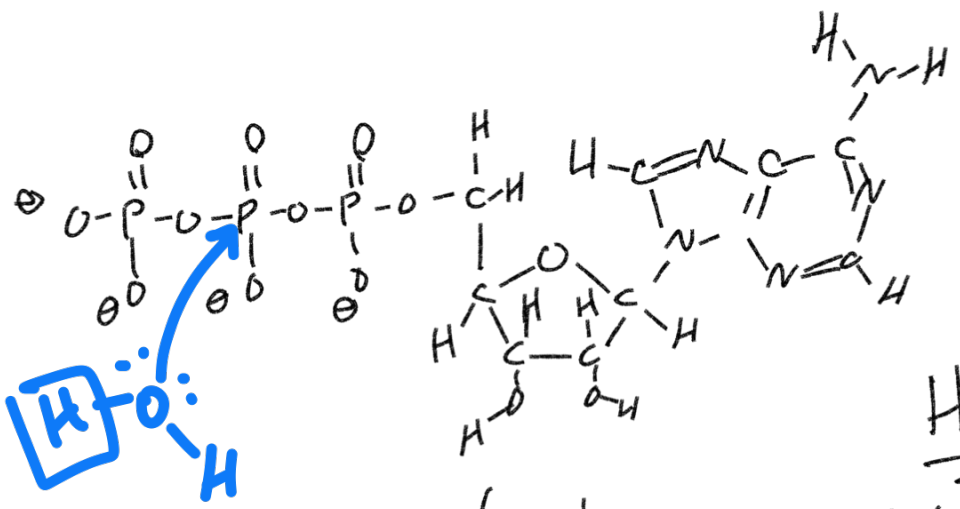


$$\Delta G = +\Delta H - T\Delta S$$

- ΔG: change in Gibbs free energy
- $+\Delta H$: change in Enthalpy (bond energy)
- T: absolute temp
- ΔS: change in entropy "chaos" degree freedom

ATP - Adenosine Triphosphate

+5



energy currency

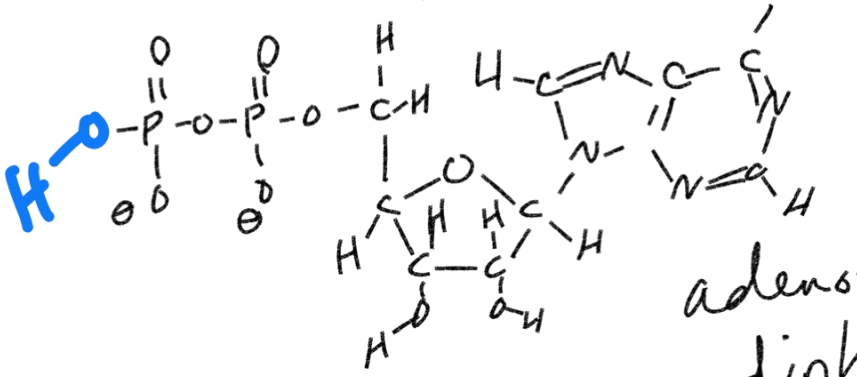
Hydrolysis

$\Delta G = -32\ kJ/mol$ ↓

exergonic

catabolic → releases energy

Favorable spontaneously



ADP

adenosine diphosphate

Catabolic Pathway Big → Lil
→ energy released

- **Catabolic pathways** release energy by breaking down complex molecules into simpler compounds
  - Eg. digestive enzymes break down food → release energy

- **Anabolic pathways** consume energy to build complex molecules from simpler ones
  - Eg. amino acids link to form muscle protein

Anabolic Lil → Big
absorb energy

ENERGY = CAPACITY TO DO WORK

- **Kinetic energy (KE)**: energy associated with motion

$KE = \frac{1}{2}mv^2$

  - *Heat* (thermal energy) is KE associated with random movement of atoms or molecules

- **Potential energy (PE)**: stored energy as a result of its position or structure

  - *Chemical energy* is PE available for release in a chemical reaction

- Energy can be converted from one form to another

  - Eg. chemical → mechanical → electrical

THE FIRST LAW OF THERMODYNAMICS (CONSERVATION OF ENERGY)

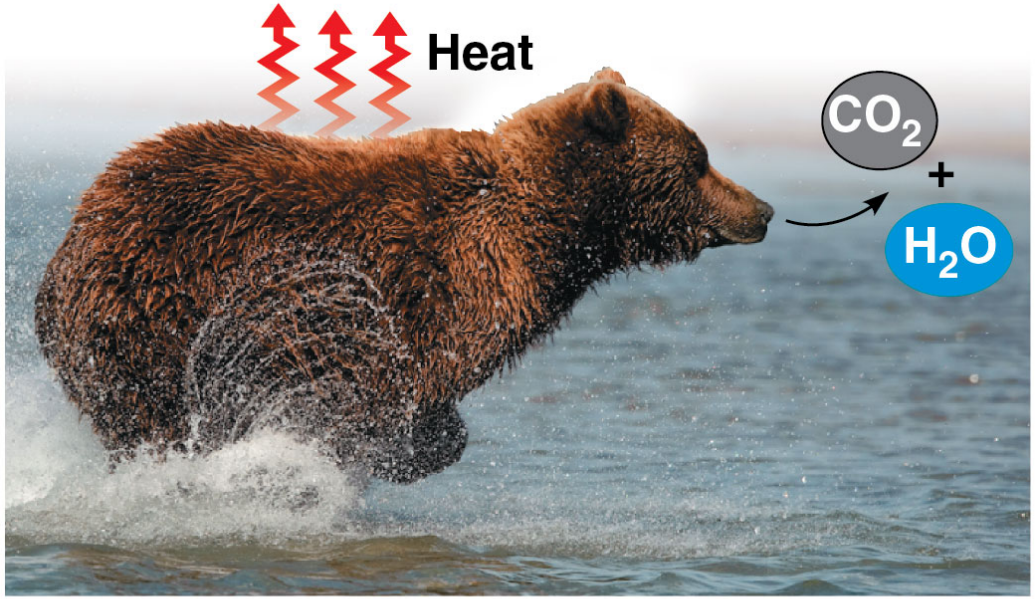
- The energy of the universe is constant
  - Energy **can** be transferred and transformed
  - Energy **cannot** be created or destroyed

(a) First law of thermodynamics

THE SECOND LAW OF THERMODYNAMICS

- Every energy transfer or transformation increases the entropy (disorder) of the universe

- During every energy transfer or transformation, some energy is *unusable*, often lost as **heat**



(b) Second law of thermodynamics

Free energy → energy available to do work Work = F·d

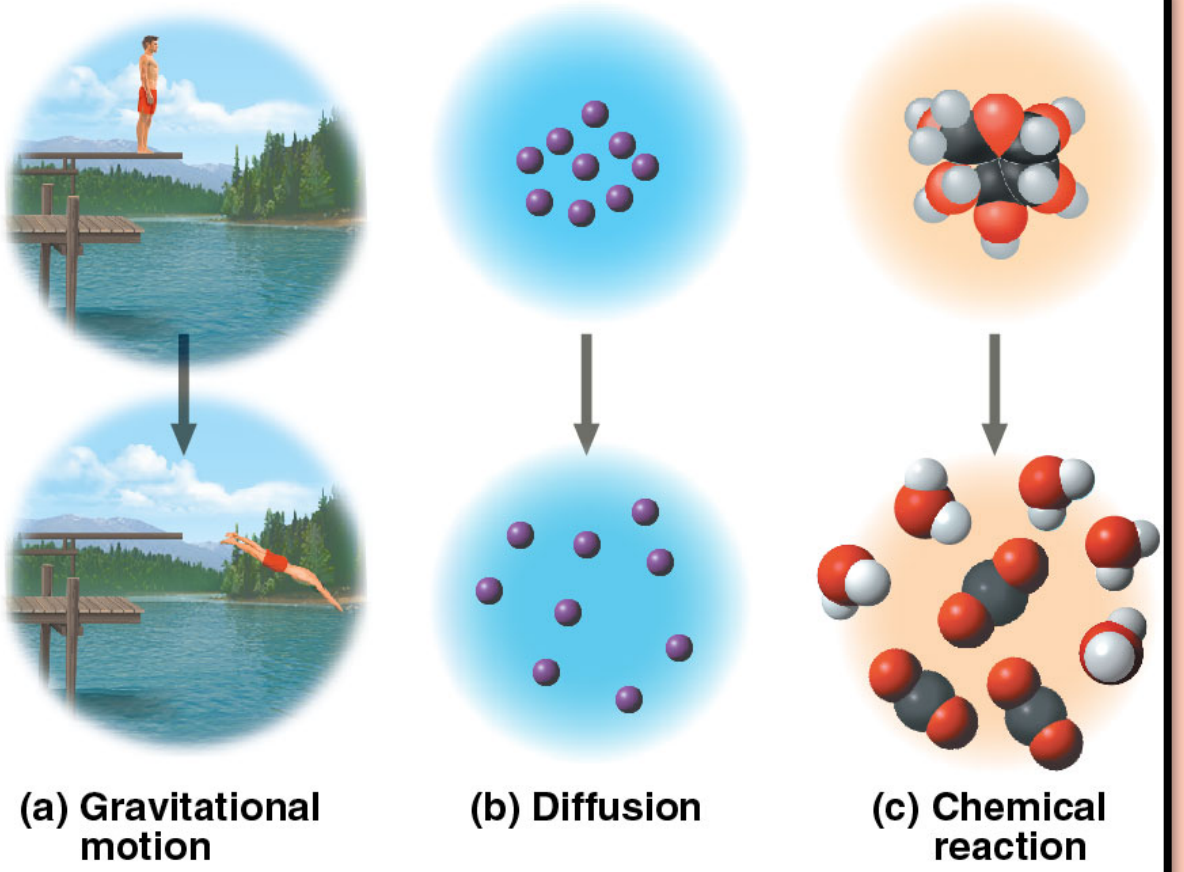
Gibbs free energy

- **Free energy**: part of a system's energy available to perform work
  - ΔG = change in free energy
- **Exergonic reaction**: energy is released
  - Spontaneous reaction often catabolic
  - $\Delta G < 0$ $-\Delta G$
- **Endergonic reaction**: energy is required
  - Absorb free energy
  - $\Delta G > 0$

- More free energy (higher *G*)
- Less stable
- Greater work capacity

In a spontaneous change
- The free energy of the system decreases ($\Delta G < 0$)
- The system becomes more stable
- The released free energy can be harnessed to do work

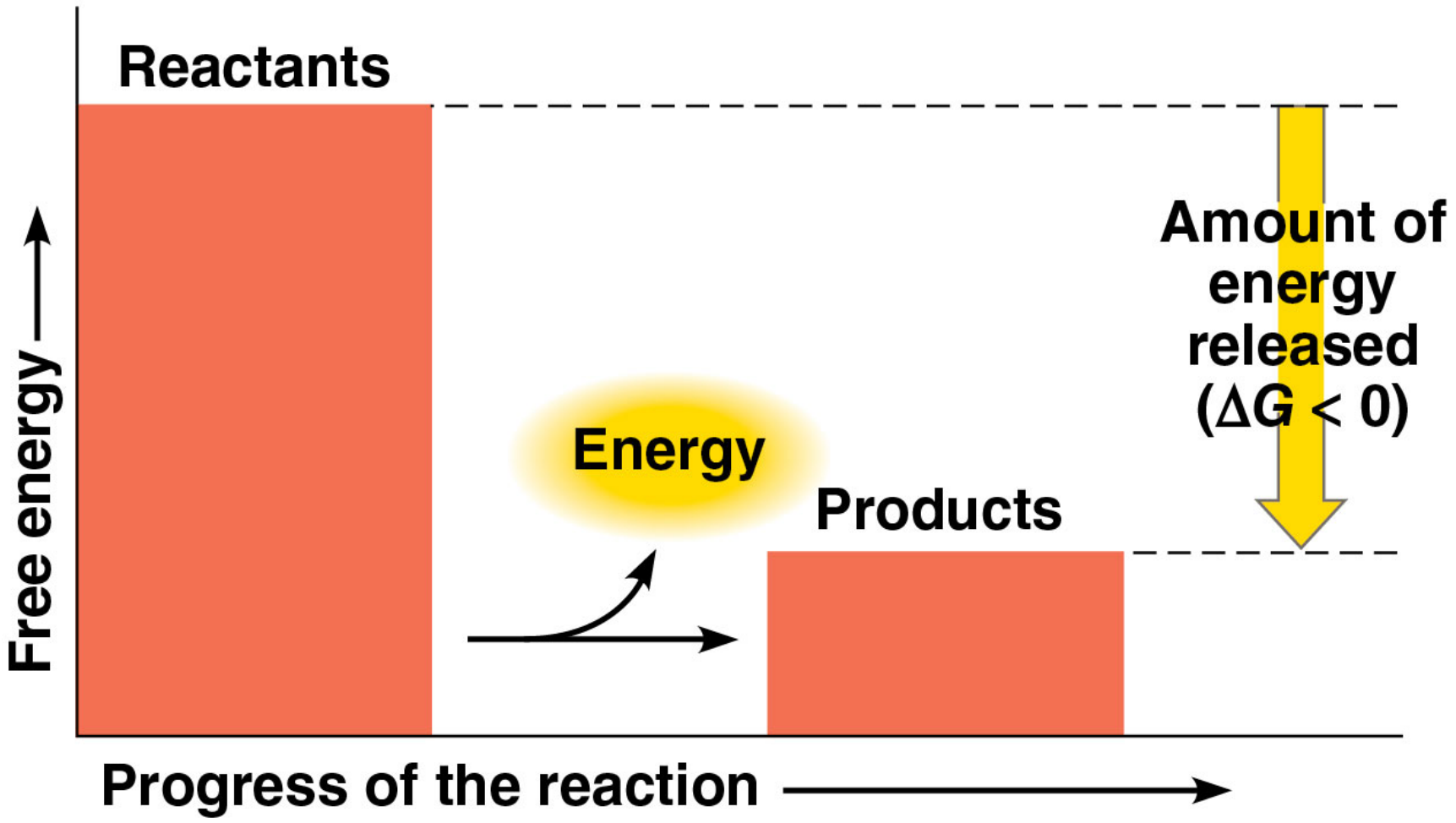
- Less free energy (lower *G*)
- More stable
- Less work capacity

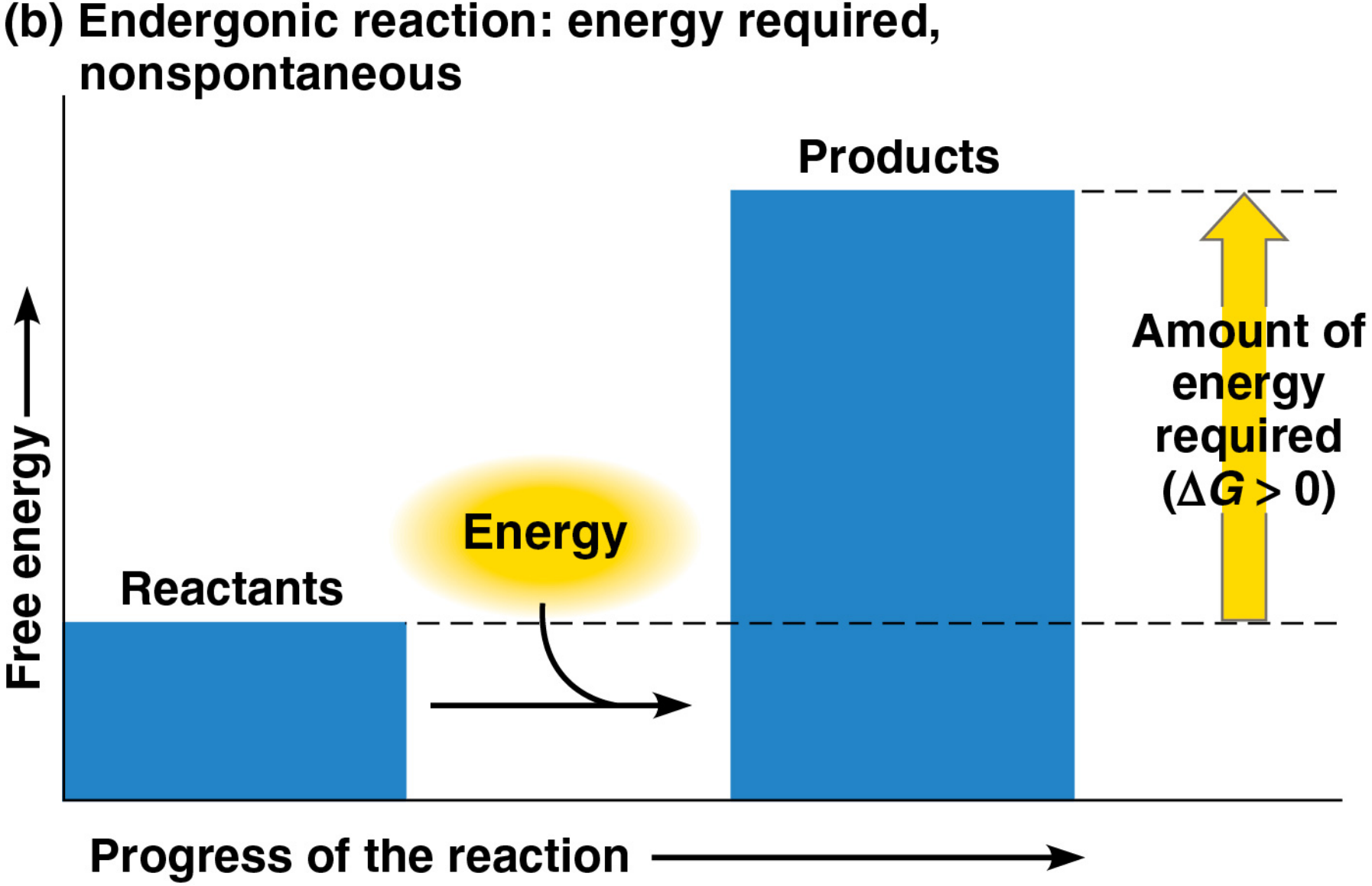


(a) Gravitational motion

(b) Diffusion

(c) Chemical reaction



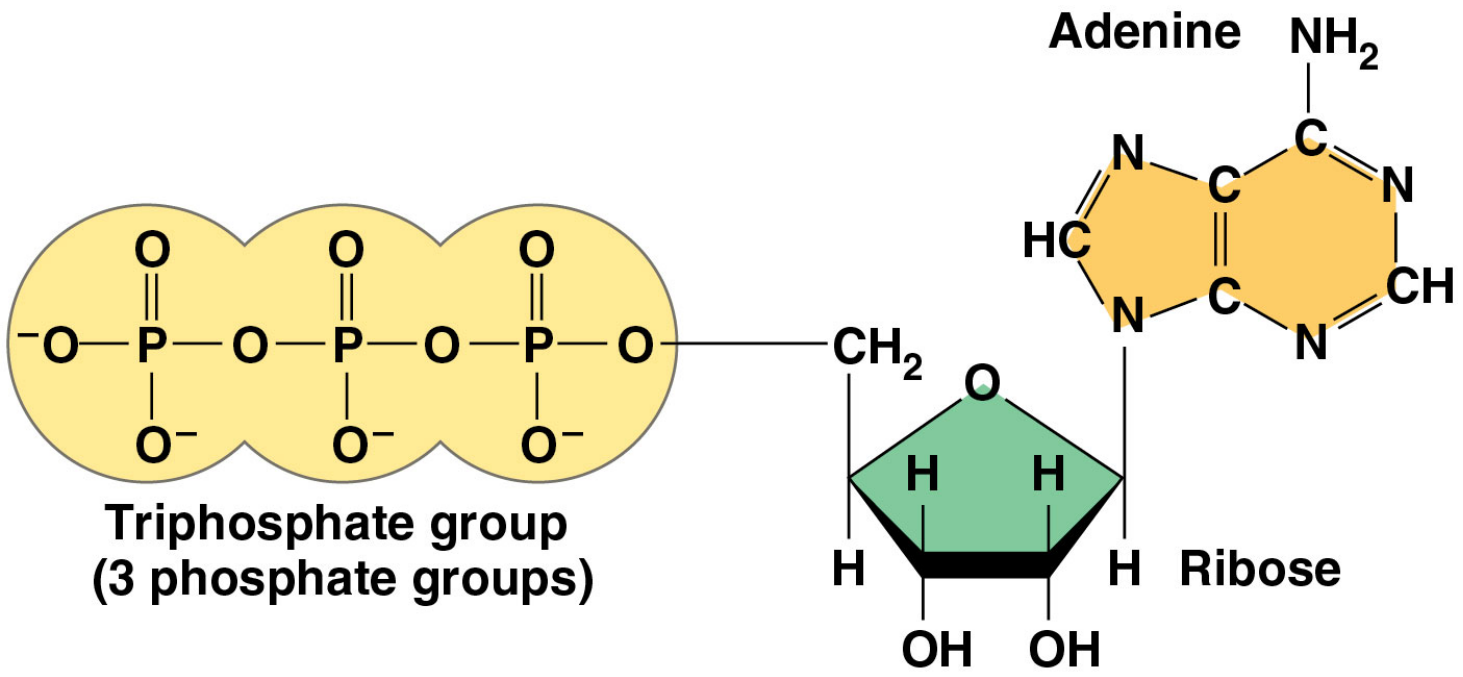


- A living cell is NOT at equilibrium
  - Constant flow of materials in/out of cell

- A cell does three main kinds of work:
  1. Mechanical
  2. Transport
  3. Chemical

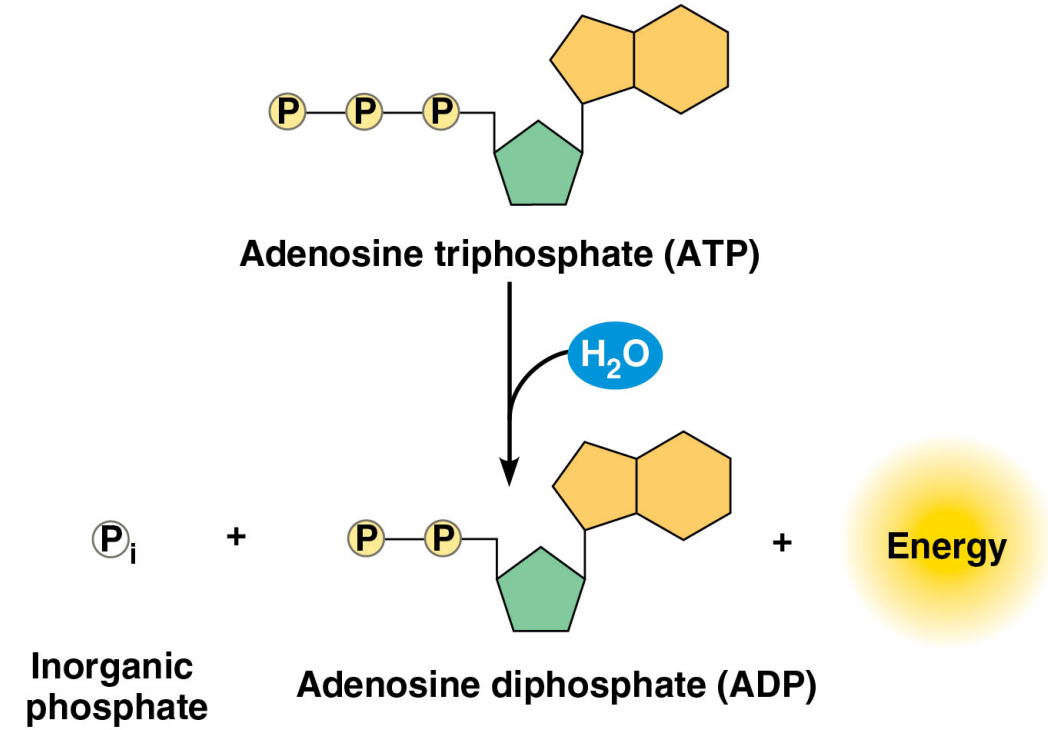
- Cells manage energy resources to do work by energy coupling: using an ***exergonic*** process to drive an ***endergonic*** one

- **ATP** (**adenosine triphosphate**) is the cell's main energy source in energy coupling
- ATP = adenine + ribose + 3 phosphates



(a) The structure of ATP

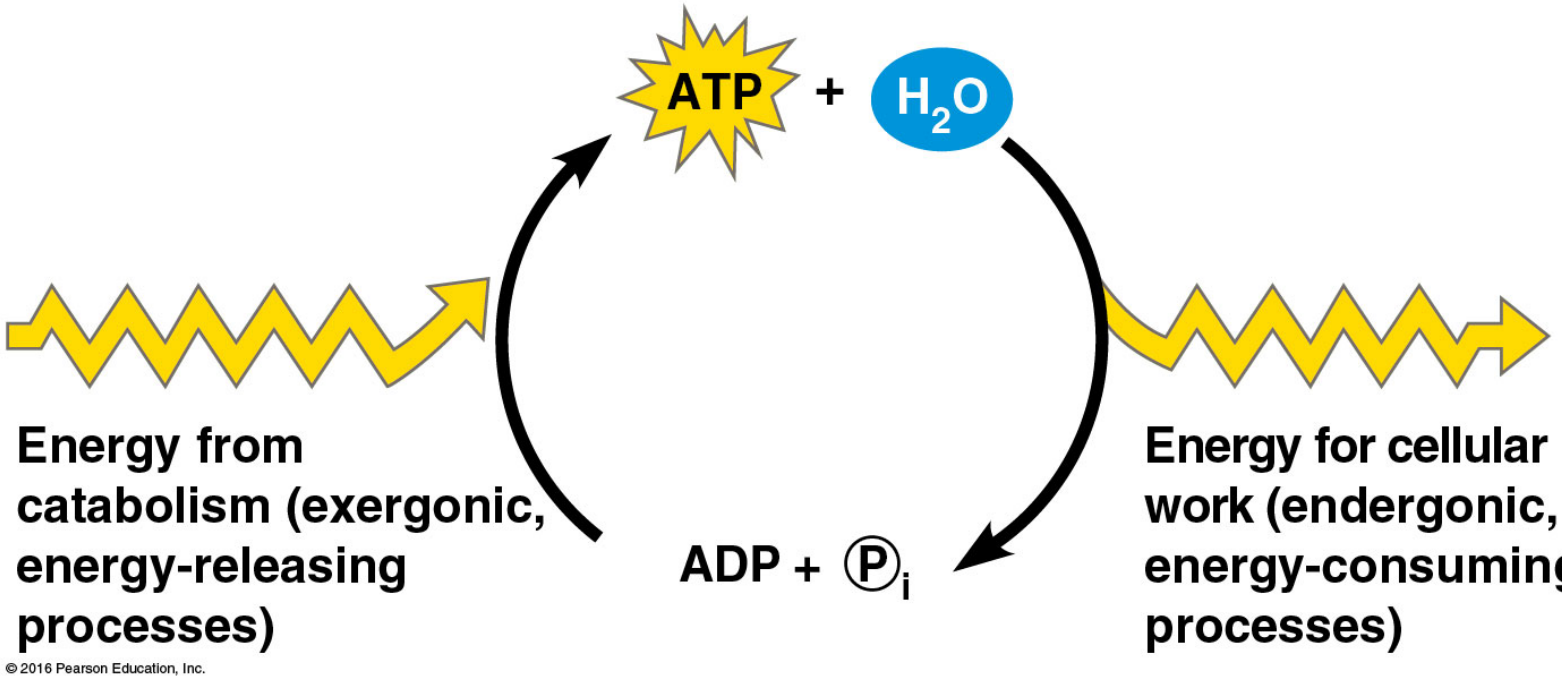
- When the bonds between the phosphate groups are broken by hydrolysis → ***Energy is released***
- This release of energy comes from the chemical change to a state of lower free energy, not in the phosphate bonds themselves

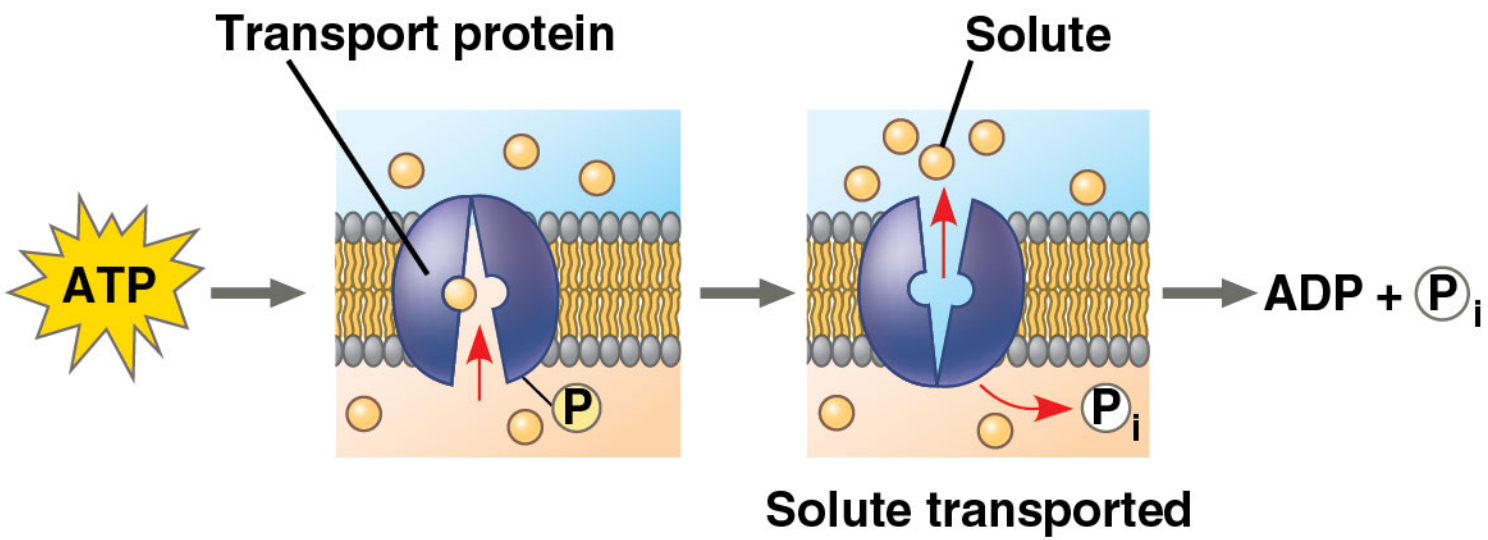


(b) The hydrolysis of ATP

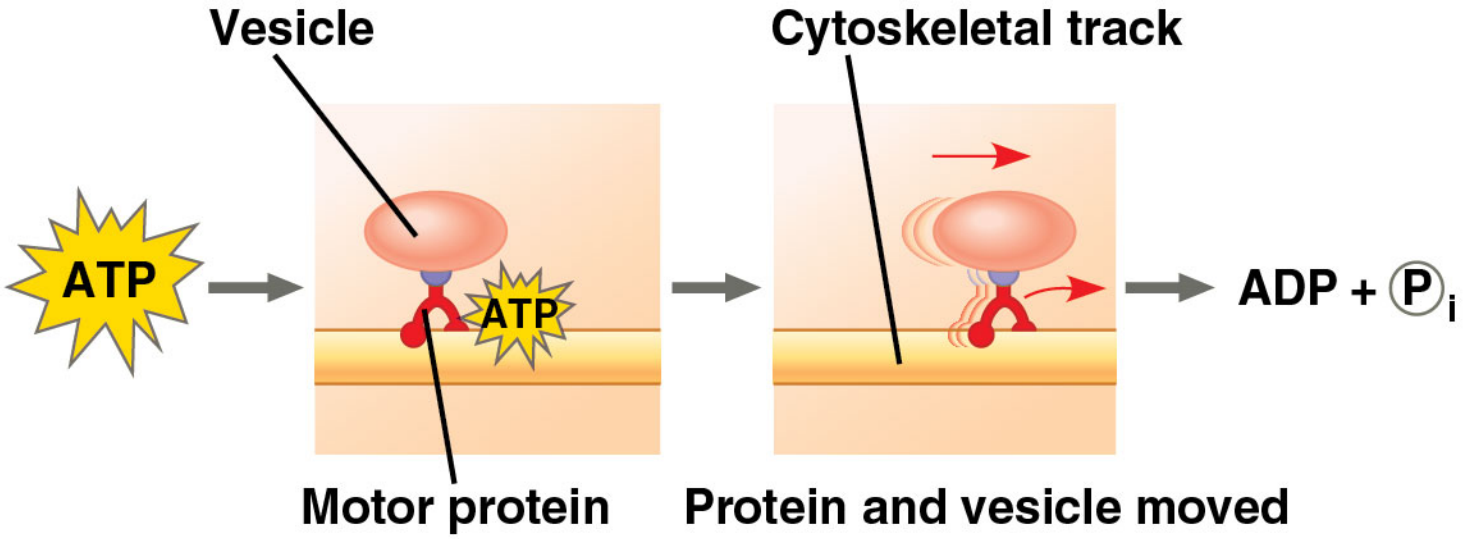
HOW ATP PERFORMS WORK

- ***Exergonic*** release of P_i is used to do the ***endergonic*** work of cell
- When ATP is hydrolyzed, it becomes ADP (adenosine diphosphate)



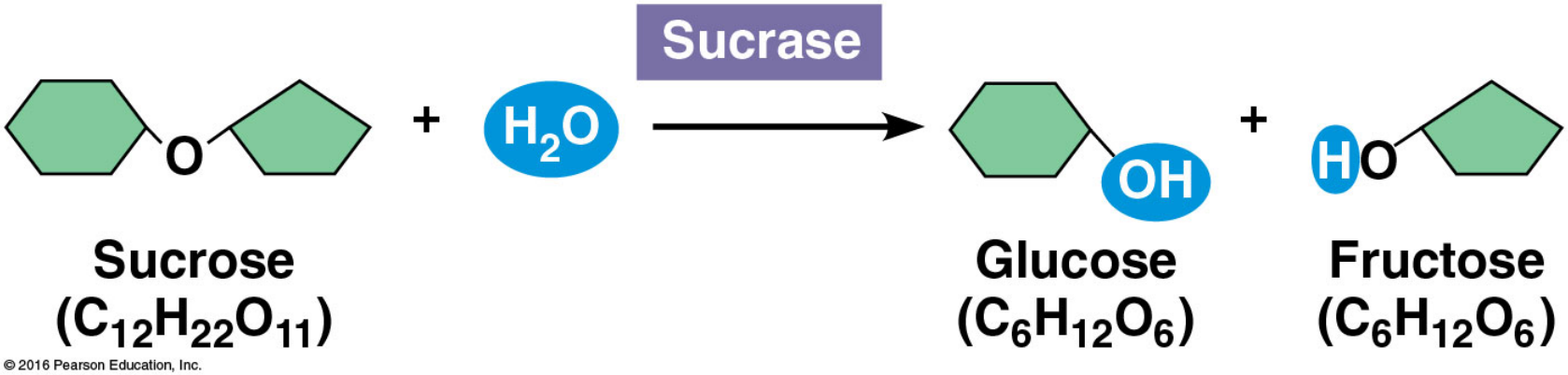


(a) Transport work: ATP phosphorylates transport proteins.



(b) Mechanical work: ATP binds noncovalently to motor proteins and then is hydrolyzed.

- Catalyst: substance that can change the rate of a reaction without being altered in the process
- Enzyme = biological catalyst



- breaking bonds)