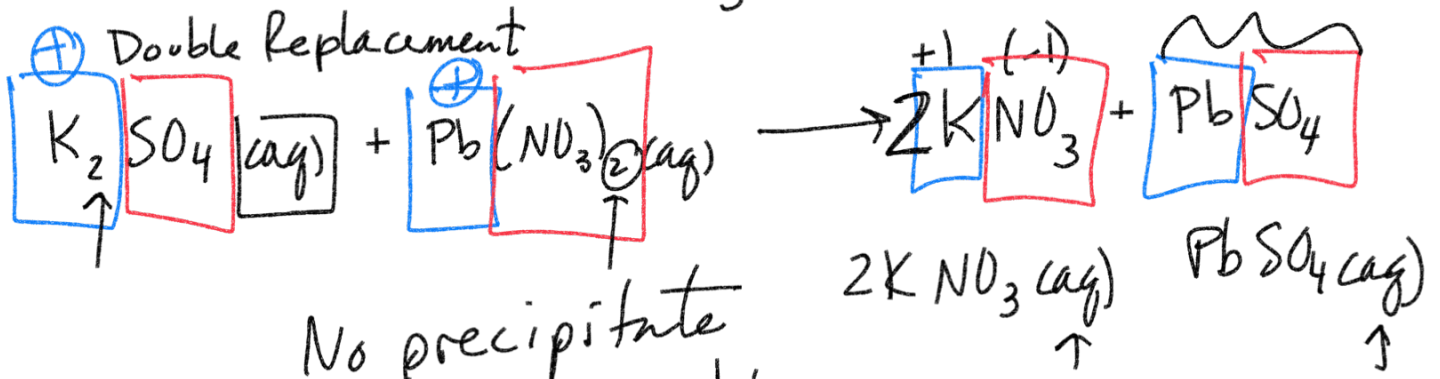
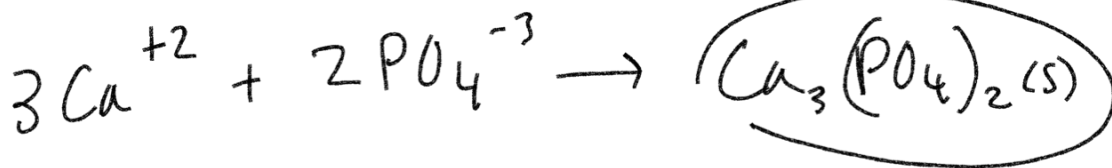
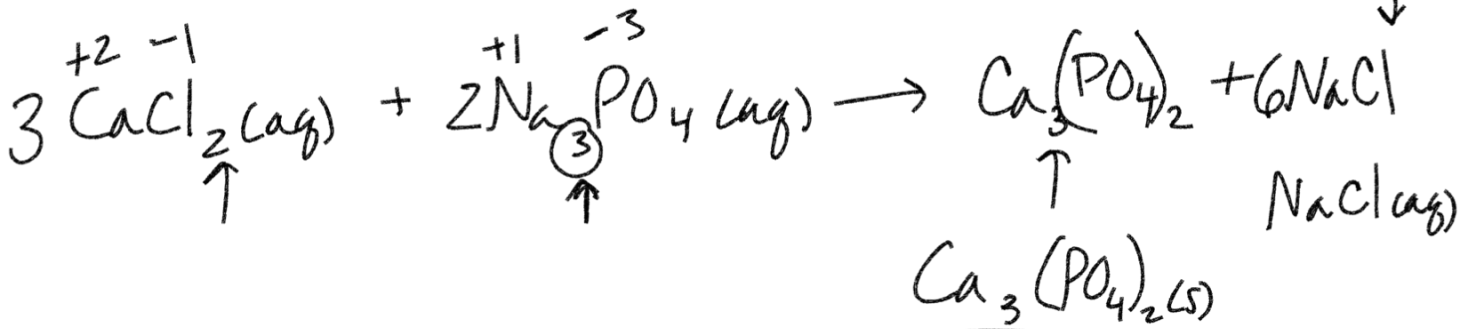


⊕ Double Replacement



No precipitate
Nothing is insoluble

NR



Copper (II) Nitrate



0.1 M

Molar Mass

$$\text{Cu} \rightarrow 63.546 \text{ g} \rightarrow 63.546 \text{ g}$$

$$\text{N} : 2 * 14.007 \text{ g} \rightarrow 28.014 \text{ g}$$

$$\text{O} : 6 * 15.999 \text{ g} \rightarrow 95.994 \text{ g}$$

$$187.554 \text{ g}$$

0.1 M =

$$M = \frac{\text{mol}}{\text{L}}$$

0.1 mol *

$$\frac{187.554 \text{ g}}{1 \text{ mol}}$$

$$= 18.7554 \text{ g}$$

↓ introduce
to
1 Liter of water