TH-GC General Chemistry Week 21 2/23

Double Replacement

$K_2SO_4(aq) + Pb(NO_3)_2(aq) \rightarrow 2KNO_3 + PbSO_4$

$2KNO_3(aq)$ $PbSO_4(aq)$

No precipitate

Nothing is insoluble

NR

$3CaCl_2(aq) + 2Na_3PO_4(aq) \rightarrow Ca_3(PO_4)_2 + 6NaCl$

$NaCl(aq)$

$Ca_3(PO_4)_2(s)$

$3Ca^{+2} + 2PO_4^{-3} \rightarrow Ca_3(PO_4)_2(s)$

Copper (II) Nitrate $Cu(NO_3)_2$

0.1 M

Molar Mass

Cu → 63.546g → 63.546g

N: 2 * 14.007g → 28.014g

O: 6 * 15.999g → 95.994g

+

187.554g

$0.1\ M = \quad M = \frac{mol}{L}$

$0.1\ mol * \frac{187.554g}{1\ mol} =$ 18.7554g

↓ introduce to 1 Liter of water